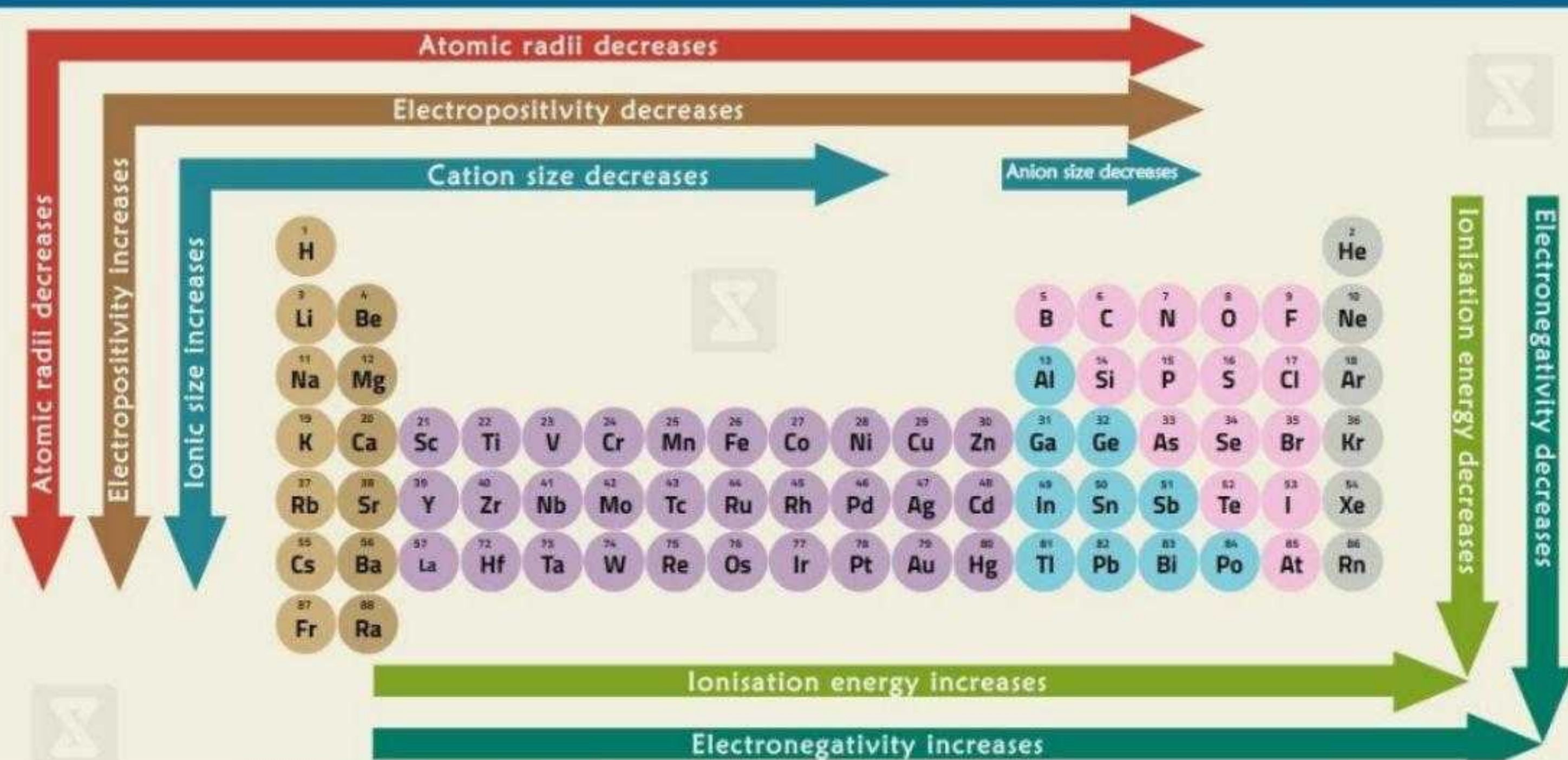


PERIODIC TABLE TRENDS



Atomic Radius

Along period, extra orbit is added therefore size of atom increases.
Along group, no. of protons increases therefore size decreases.

Largest atom

Cs

Electropositivity

Ability of a metal to lose electrons increases with increase in its size.

Largest electropositive atom

Cs

Ionic Size

Ionic size depends on $\frac{e}{p}$ ratio.
As protons increase, attraction over electron increases therefore size of ion decreases.

Ionisation Energy

As size increases it becomes easy to remove outer electron. Thus ionisation energy decreases along the group

Largest ionisation energy required

F

Electronegativity

Ability of an atom to accept electron increases with decrease in its size.

Largest electronegative atom

F